## In-situ TEM Study on Electrochemical Behavior of α-MnO<sub>2</sub> Nanowire

Yifei Yuan<sup>1</sup>, Anmin Nie<sup>2</sup>, Sunand Santhanagopalan<sup>3</sup>, Dennis Desheng Meng<sup>3</sup>, Reza Shahbazian-Yassar<sup>2\*</sup>

<sup>2</sup> Department of Mechanical Engineering-Engineering Mechanics, Michigan Technological University, 1400 Townsend Dive, Houghton, Michigan 49931, United States

Manganese dioxide (MnO<sub>2</sub>) is widely known to possess various allotropic forms such as  $\alpha$ -,  $\beta$ - and  $\gamma$ -phases, which are constructed by combination of octahedral [MnO<sub>6</sub>] building blocks to form different tunneled structures. These special structures are believed to account for the various characteristics of MnO<sub>2</sub> when it is employed as electrode material in lithium (ion) batteries <sup>[1-2]</sup>. There is, however, lack of direct proof demonstrating the role of tunneled structure during electrochemical lithiation/delithiation of MnO<sub>2</sub>.

In this work, by applying high resolution transmission electron microscopy (HRTEM) to single  $\alpha$ -MnO<sub>2</sub> nanowire along both axial and radial directions, the tunneled structure is clearly shown and characterized. The  $\alpha$ -MnO<sub>2</sub> nanowire is proved to be single crystalline and grow along [001] direction. Cross-sectional HRTEM images have shown that the nanowire has a squared cross section and 2x2 tunnels align parallelly along its growth direction [001], matching very well with simulated crystal structure. An in-situ TEM setup for study of MnO<sub>2</sub>'s dynamic lithiation/delithiation process is also designed and demonstrated. This open-cell design inside TEM allows real time observation of electrode behavior during its charge and discharge process, enabling better understanding of electrochemical essentials of  $\alpha$ -MnO<sub>2</sub> and possible modifications of its composition, morphology and structure to further improve its overall performance in battery application.

## References:

- [1] Chen K., Dong Noh Y., Li K., Komarneni S., Xue D., The Journal of Physical Chemistry C 117 (2013), p.10770.
- [2] Xun Wang, Y. L., Journal of the American Chemical Society 124 (2002), p.2.

<sup>&</sup>lt;sup>1</sup> Department of Materials Science and Engineering, Michigan Technological University, 1400 Townsend Dive, Houghton, Michigan 49931, United States

<sup>&</sup>lt;sup>3.</sup> Department of Mechanical and Aerospace Engineering, University of Texas at Arlington, 500 West 1<sup>st</sup> Street, Arlington, Texas 76019, United States (Part of the work has been done at Michigan Technological University<sup>2</sup>)

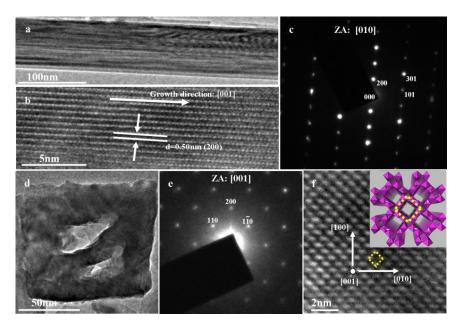


Figure 1 a: TEM image of single  $\alpha$ -MnO<sub>2</sub> nanowire; b: HRTEM of  $\alpha$ -MnO<sub>2</sub> nanowire along [010] direction; c: Electron diffraction pattern of  $\alpha$ -MnO<sub>2</sub> nanowire along [010] zone axis; d: Cross-sectional image of  $\alpha$ -MnO<sub>2</sub> nanowire; e: Electron diffraction pattern along [001] zone axis; f: Cross-sectional HRTEM and inserted simulation result along [001] direction showing 2X2 tunnels indicated by yellow dashed squares.

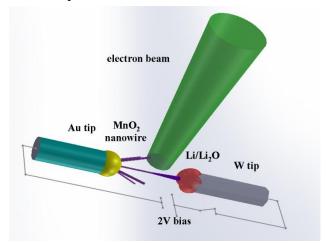


Figure 2 Cartoon for the in-situ TEM setup testing  $\alpha$ -MnO<sub>2</sub> nanowire's lithiation/delithiation behavior. The nanowire is attached to a gold tip by conductive epoxy, while Li metal is attached to a tungsten tip and functions as the counter electrode. Lithiation starts when a constant potential of -2 V is applied to MnO<sub>2</sub> against Li counter electrode (+4 V for the delithiation process).